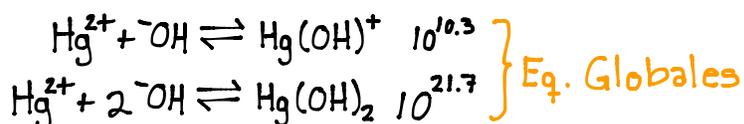
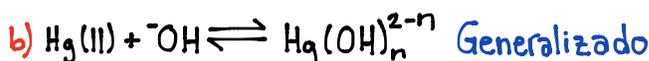
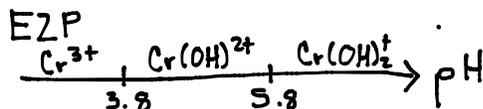
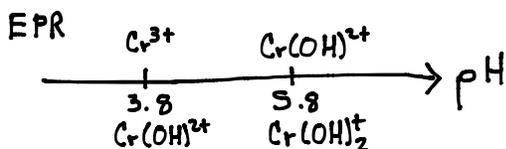
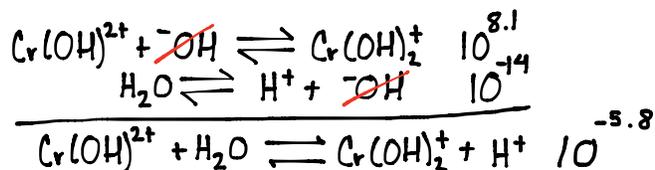
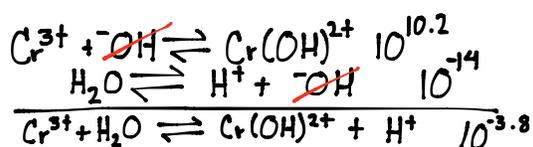
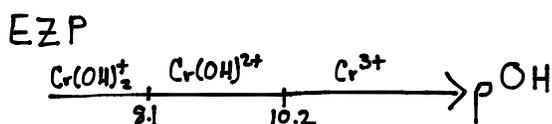
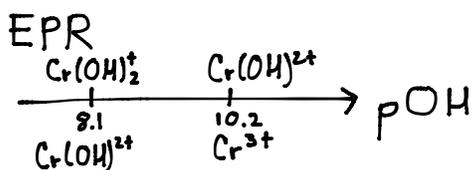
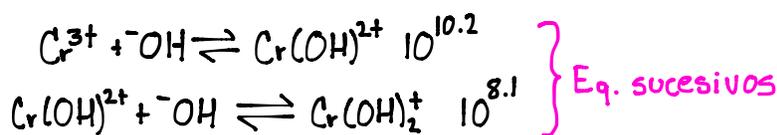
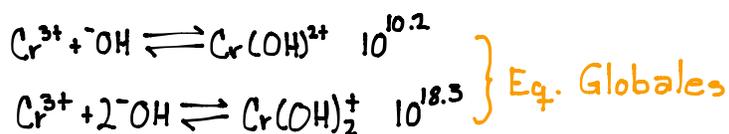
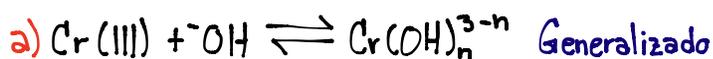
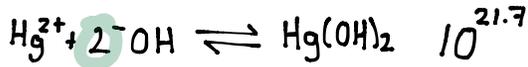
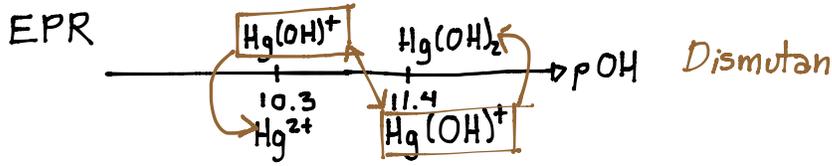
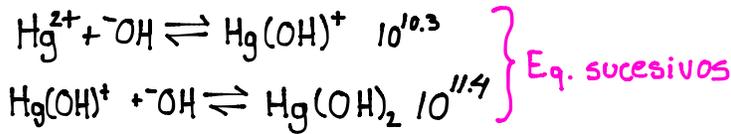


Establecer los equilibrios parciales de formación calculando las constantes de equilibrio respectivas, determinar la escala de predicción de reacciones, obtener la E.Z.P. Y en los casos que sea posible obtener la escala en función del pH para los siguientes complejos (uno por inciso):

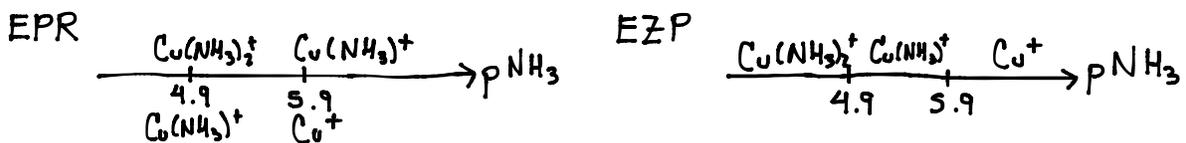
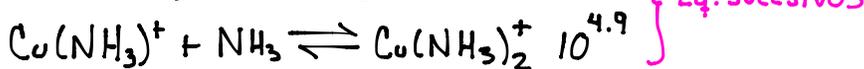
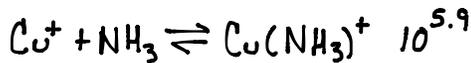
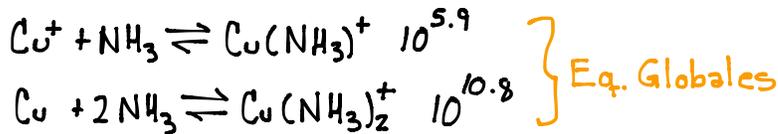
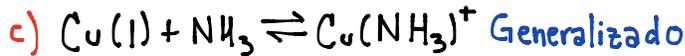
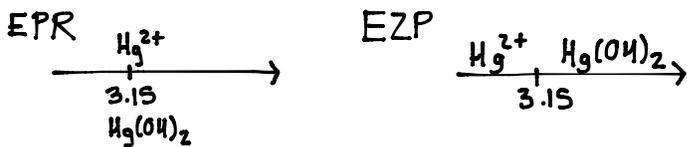
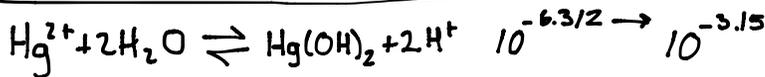
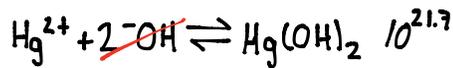
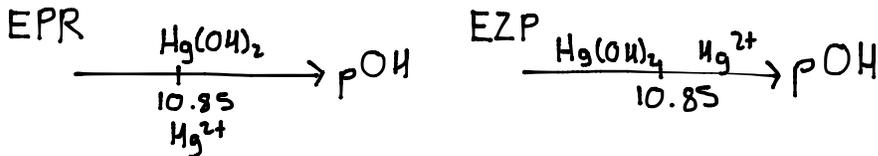
	$\log \beta_1$	$\log \beta_2$
a) $\text{Cr}(\text{OH})_n^{3-n}$	10.2	18.3
b) $\text{Hg}(\text{OH})_n^{2-n}$	10.3	21.7
c) $\text{Cu}(\text{NH}_3)_n^{1-n}$	5.9	10.8
d) $\text{Au}(\text{SCN})_n^{1-n}$	-	2.5

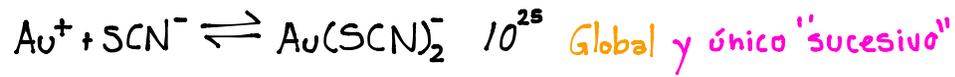
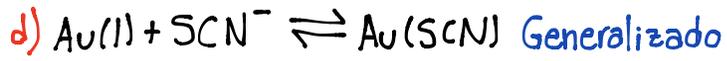




$$K_c = 10^{\frac{21.7}{2}} = 10^{10.85}$$

↓
Dos partículas





$2p_{SCN} = 25$

$p_{SCN} = \frac{25}{2}$
 $p_{SCN} = 12.5$

